

SOLUTIONS

CLASSIC

Molarity

Date:

- 20 g NaOH is dissolved in water to form 200 ml solution. What is the molarity of the solution?(2.5 M) ($MM_{\text{NaOH}}=40$ g/mole)
- 50 g CaBr_2 is dissolved in water to form 500 ml solution. What is the molarity of the solution?(0.5 M) ($MM_{\text{CaBr}_2}=200$ g/mole)
- 48 g CH_3OH is dissolved in water to form 1500 ml solution. What is the molarity of the solution?(1 M) ($MM_{\text{CH}_3\text{OH}}=32$ g/mole)
- 126 g HNO_3 is dissolved in water to form 200 ml solution. What is the molarity of the solution?(10 M) ($MM_{\text{HNO}_3}=63$ g/mole)
- 28 g KOH is dissolved with enough amount of water to form 2000 ml solution. What is the molar concentration of the solution? (2.5 M)
- What is the molar concentration of 1200 ml solution that contains 60 g NaOH? (1.25 M)
- What is the mass of HNO_3 that is in 2 L of 1 M HNO_3 solution? (126 g)
- What is the mass of CaBr_2 that is in 500 ml of 1.5 M CaBr_2 solution? (150 g)
- What is the volume of 3 M NaOH solution that contains 360 g NaOH? (3 L)
- What is the volume of 2 M CH_3OH solution that contains 16 g CH_3OH ?(0.25 L)
- What is the molarity of 50% HCl solution by mass that has a density of 1.17 g/ml?(16 M) ($MM_{\text{HCl}}=36.5$ g/mole)
- What is the molarity of 10% AgNO_3 solution by mass that has a density of 1.17 g/ml?(0.7 M) ($MM_{\text{AgNO}_3}=170$ g/mole)
- 400 g of 16% CH_3OH solution by mass is dissolved with enough amount of water and 4000 ml solution is formed. What is the molarity of the new solution? (0.5 M)
- 40 ml of CH_3OH ($d=0.9$ g/ml) is dissolved in water and 500 ml solution is obtained. What is the molarity of the new solution? (2.25 M)
- What is the molarity of 48% HBr solution by mass that has a density of 1.50 g/ml?(8.88 M) ($MM_{\text{HBr}}=81$ g/mole)
- What is the molarity of 47% HI solution by mass that has a density of 1.50 g/ml?(5.51 M) ($MM_{\text{HI}}=128$ g/mole)
- What is the molarity of 14% KOH solution by mass that has a density of 1.2 g/ml? (3 M)
- What is the molarity of 20% NaOH solution by mass that has a density of 1.2 g/ml? (6 M)
- What is the volume of 4M solution that contains 8 g NaOH? (50 ml)
- What is the mass of CH_3OH that is in 750 ml of 1.5 M CH_3OH solution?(36 g) ($MM_{\text{CH}_3\text{OH}}=32$ g/mole)
- What is the mass of HCl that is in 250 ml of 2 M HCl solution?(18.25 g) ($MM_{\text{HCl}}=36.5$ g/mole)
- What is the volume of 1.5 M HCl solution that contains 36.5 g HCl?(0.66 L) ($MM_{\text{HCl}}=36.5$ g/mole)
- What is the volume of 0.75 M NaOH solution that contains 120 g NaOH?(4 L) ($MM_{\text{NaOH}}=40$ g/mole)
- Solid NaOH is added into 200 ml of 2M NaOH solution and 200 ml of 3M NaOH solution is formed. What is the mass of added NaOH?(8 g)