

CHEMISTRY DAILY PLAN

Class:

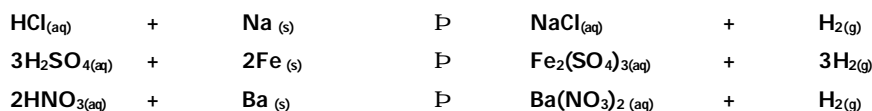
Date:

Subject: Other Types of Reactions

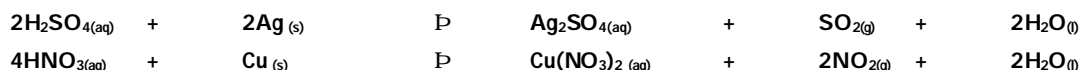
Time:

a) Acids - metals reactions.

Acids react with metals to produce hydrogen gas. (Except: Cu, Hg, Ag, Pt and Au)

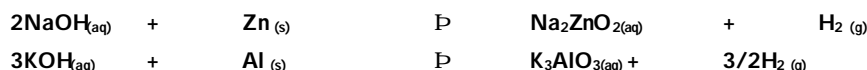


Copper, mercury and silver react only sulfuric acid and nitric acid to produce SO₂ and NO₂.



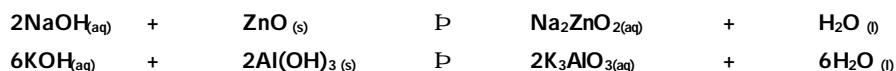
b) Bases - metals reactions.

Generally metals do not react with bases. But amphoteric metals which are Al, Zn, Cr, Sn and Pb react with bases to produce H₂ gas.



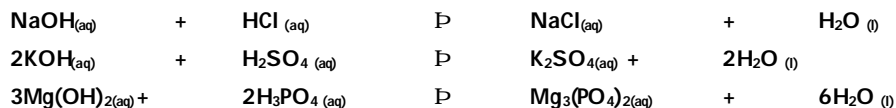
Na₂ZnO₂
Sodium zincate
K₃AlO₃
Potassium aluminate

Same reactions occur with amphoteric oxide and amphoteric hydroxide. But instead of hydrogen, water will produce.

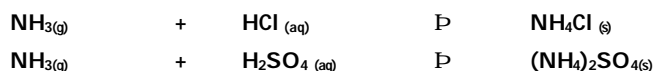


c) Neutralisation reactions.

The reaction of acids with bases is called neutralisation reactions. From neutralisation reactions water and salt will produced.



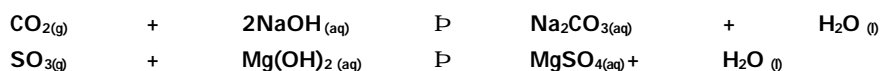
NH₃ is base with no OH. If this base reacts with acids only a salt will produced.



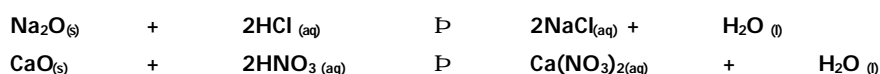
If acids react with carbonates which are salts with basic characteristic. Salts, water and carbon dioxide gas will produced.



From the reaction of acidic oxide and base, salt and water form.



From the reaction of basic oxide and acid, salt and water form.



From the reaction of acidic oxide with basic oxide only salt form.

