

CHEMISTRY DAILY PLAN

Class:

Date:

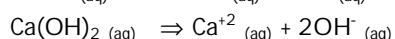
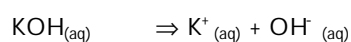
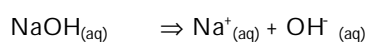
Subject: Bases

Time:

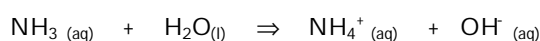
Some substances in our daily life have basic characteristics. Toothpaste, soap, washing soda, ammonia solution, are good examples of bases.

Definition

The compounds that produces OH⁻ ions in water are called bases, Ionization reactions of some bases are as follow;



The ionization reaction of anhydrous base, NH₃ is;



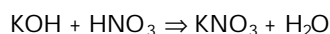
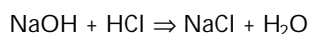
Naming Bases

In naming bases, we say "hydroxide" after the name of the metals.

<u>Formula</u>	<u>Name</u>	<u>Formula</u>	<u>Name</u>
NaOH	Sodium hydroxide	Al(OH) ₃	Aluminum hydroxide
AgOH	Silver hydroxide	Ca(OH) ₂	Calcium hydroxide
Mg(OH) ₂	Magnesium hydroxide	Fe(OH) ₂	Iron (II) hydroxide
Cu(OH) ₂	Copper (II) hydroxide	KOH	Potassium hydroxide

General properties of bases

1. Bases have bitter taste.
2. They have slippery feeling. Remember the slippery feeling of soaps, which are mildly basic substances.
3. They form conducting solutions, because they are ionized.
4. They change the color of red litmus paper to blue.
5. Bases react with acids to produce salt and water (Neutralization Reactions).



4. Strength of bases

NaOH and KOH are strong bases that completely ionize in water and their aqueous solutions are good conductors of electricity. Bases like NH₃, AgOH, Cu(OH)₂, Al(OH)₃ are weak bases that partially ionize in water, and they are poor conductors of electricity.

