

ACIDS AND BASES

TEST

pH and pOH – 1

Date:

- Information for three solutions are:
I. pH = 7
II. pOH = 12
III. $[H^+] = 1 \times 10^{-4}$
Which one(s) of the above is acidic solution?
A) I – II B) I – III C) II D) III E) II – III
- What is the pH of the solution if $[H^+] = 1 \times 10^{-4}$.
A) 2 B) 4 C) 8 D) 10 E) 12
- What is the pH of the solution if $[OH^-] = 1 \times 10^{-2}$.
A) 2 B) 4 C) 8 D) 10 E) 12
- What is the $[H^+]$ of the solution if pH = 7.
A) 1 B) 10 C) 10^{-3} D) 10^{-7} E) 10^{-14}
- What is the pOH of neutral solution?
A) 0 B) 1 C) 7 D) 10 E) 14
- What is the pH of 0.005 M H_2SO_4 solution?
A) 1 B) 2 C) 5 D) 10 E) 12
- What is the pOH of 0.001 M NaOH solution?
A) 1 B) 3 C) 5 D) 7 E) 11
- What is the pOH of 5×10^{-5} M $Mg(OH)_2$ solution?
A) 2 B) 4 C) 8 D) 10 E) 11
- What is the pH of 0.01 M NaOH solution?
A) 2 B) 6 C) 8 D) 9 E) 12
- What is the pOH of 0.01 M HNO_3 solution?
A) 2 B) 4 C) 8 D) 10 E) 12
- What is the pOH of 0.05 M H_2SO_4 solution?
A) 1 B) 7 C) 9 D) 10 E) 13
- What is the pH of 5×10^{-4} M $Ca(OH)_2$ solution?
A) 2 B) 3 C) 8 D) 10 E) 11
- What is the pH of the solution that is prepared from 7.3 g HCl with 20 L water? (HCl:36.5)
A) 1 B) 2 C) 4 D) 8 E) 12
- What is the pH of the solution that is prepared from 0.49 g H_2SO_4 with 10 L water? ($H_2SO_4 = 98$)
A) 1 B) 3 C) 5 D) 8 E) 11
- What is the pH of the solution that is prepared from 4 g NaOH with 10 L water? (NaOH:40)
A) 1 B) 2 C) 4 D) 8 E) 12
- What is the pOH of 1 L of 0.5 M $Ba(OH)_2$ solution after addition 9 L water?
A) 1 B) 3 C) 5 D) 8 E) 11